

## OBJECT OF THE INVENTION

The process object of the patent entails obtaining electrolytic manganese from treated  
 sludge of the exhaust gases of ferroalloy production furnaces for any other industrial  
 5 waste having magnesium in general. The intention is to achieve the integrated  
 management of a waste which is produced worldwide in significant amounts by  
 making use of its manganese content, which is the primary raw material in the actual  
 production process in which it is generated.

## BACKGROUND OF THE INVENTION

Manganese is a metal which is not found in nature as a free form, but rather combined  
 in minerals in the composition of which it is found as an oxide (primarily pyrolusite)  
 or carbonate (primarily rhodochrosite). Production processes traditionally requiring  
 15 the addition of manganese have been carried out using these minerals in iron  
 metallurgy reduction furnaces for the purpose of using the manganese as an alloy  
 element or as a deoxidizing and desulphurizing agent in steel production. However, at  
 the end of the 19th century, different methodologies for obtaining pure manganese  
 from manganese minerals began to be studied for the purpose of improving the alloys  
 20 in which it participated and for broadening its field of application, several processes  
 being developed: aqueous electrolysis of manganese salts, electrothermic,  
 carbothermic, aluminothermic and silicothermic processes.

Among the background, US patents 2,511,507 for "TREATING MANGANESE  
 ELECTROPLANTING SOLUTIONS"; 2,483,287 for "METHOD OF PURIFYING  
 25 MANGANESE ELECTROLYTES"; 2,347,451 "ELECTROLYTIC DEPOSITION  
 OF MANGANESE"; 2,343,293 "PROCESS FOR THE PURIFICATION OF  
 MANGANOUS SULPHATE SOLUTIONS "; and U.K. patent 528.112  
 "IMPROVEMENTS IN THE ELECTROLYTIC PRODUCTION OF  
 MANGANESE" are worth mentioning.

Of all of these processes, the one which allows obtaining greater manganese purity, has very accessible production costs and, therefore, is the most used process, is the first one, i.e. the production of manganese by means of aqueous electrolysis of salts of this metal or obtaining electrolytic manganese. This product is currently marketed with purities ranging between 99.5 and 99.9% metal.

The electrolytic route for obtaining manganese metal was first researched by Davis in 1930. However, this process did not become important until 1939, when the demand for electrolytic manganese by steel producers (for the manufacture of weapons) forced the U.S. Bureau of Mines to install a pilot plant in Knoxville (Tennessee). This plant was redesigned in 1940, and in 1944 it reached a production capacity of 1,500 t/year. In Japan, electrolytic manganese production began in 1941. The U.S. Bureau of Mines built a second pilot plant in 1942 in Boulder City. For its part, The Electrolytic Manganese Corporation of Krugersdorp, in South Africa, began production in 1955. Currently, most of the electrolytic manganese consumed worldwide is produced in China and South Africa.

The process for obtaining electrolytic manganese which has been developed at a laboratory level stems from the need to solve the greatest environmental drawback derived from the manufacture of ferroalloys for any other industrial waste having magnesium in general, which is the generation of waste as a result of the processing of the production furnace exhaust fumes. The most effective manner of processing these emissions is wet gas scrubbing, such that the particles contained therein are retained in water. The subsequent processing of this water gives rise to obtaining a waste product having a high manganese content which is difficult to use as a recyclable material due to its physical nature.

The drawback described made the requesting company assign the Cátedra de Metalurgia de la Universidad de Oviedo (Spain) four years ago the task of carrying

out viability tests for the extraction of the manganese content of this waste by hydrometallurgical route for its subsequent electrolytic recovery, and to thus obtain a product of high added value. The research carried out has clearly shown that by means of certain novelties on the previously mentioned methods, it is possible to meet this challenge.

### DESCRIPTION OF THE PROCESS

The starting material used in the method for obtaining electrolytic manganese object of the invention is the waste obtained in the scrubbing of the ferroalloy production furnace exhaust gases, primarily made up of:

	%		%
<b>MnO</b>	5 – 50	<b>S</b>	0 – 10
<b>SiO<sub>2</sub></b>	5 – 40	<b>K<sub>2</sub>O</b>	0 – 8
<b>CaO</b>	2 – 14	<b>Na<sub>2</sub>O</b>	0 – 5
<b>Al<sub>2</sub>O<sub>3</sub></b>	1 – 12	<b>C</b>	2 – 15
<b>MgO</b>	1 – 11	<b>FeO</b>	0 – 5

The process which has been conceived consists of the following phases:

Sulphation

Lixiviation

Purification

Electrolysis

5        1) SULPHATION

10        The starting industrial waste is subjected to acid etching until dryness, on one hand a byproduct consisting of the non-etched materials being obtained, and on the other hand, the metal sulphates containing ions of the different metals capable of being  
lixivated that are present in the waste, primarily manganese.

2) LIXIVIATION

15        The lixiviation and filtering of that solid give rise to a liquor entraining the sulphates and other soluble substances generated during the processing in the furnace, and to obtaining a solid waste which is rendered inert by means of the washing thereof. The exhaust gases of this furnace are subjected to alkaline scrubbing.

3) PURIFICATION

20        The purification step is indispensable for the solution to be suitable for electrolysis. Small amounts of unwanted metals in the solution would cause the contamination of the deposit, or they would even prevent the electrodeposition reactions.

25        There are two types of impurities which must be removed from the liquor, according to the manner in which they are to be dealt with.

A first group is made up of those which can be separated by a simple pH control. It is about taking the solution to pH conditions which force the precipitation of unwanted

ions, following the Pourbaix diagrams (potential – pH), keeping the manganese in solution. The precipitation of practically all of the iron and aluminum,, as well as of other less problematic contaminants such as cobalt or nickel, is thereby ensured. To increase the pH of the liquor, lime is added while the pulp is stirred. The obtained precipitate is separated by means of filtration.

Lastly, the solution is passed through an active carbon filter prior to subjecting it to secondary purification.

The second group of contaminants is made up of base metals which cannot be completely removed by pH control, since the value which is necessary to reach for the precipitation thereof interferes with the manganese precipitation pH.

The metals identified in this second group are headed up by zinc, and they are nobler than manganese. Their removal is achieved by means of their precipitation in the form of sulphide at a slightly acid pH. This precipitation requires a sufficient residence time in order for the manganese sulphide which is formed to redissolve, but not excessive in order to prevent impurities which had precipitated from redissolving.

The already purified liquor is conditioned by means of the addition of a base until reaching an almost neutral pH in order to enable its introduction in the electrolysis cell. Lastly, part of this liquor is passed through a crystallizer to remove part of its calcium and magnesium content as ammonium salts.

#### 4) ELECTROLYSIS

Prior to being introduced in the cathode cells of special electrolysis tanks, the conditioning of the liquor is completed with the following additives:

Ammonium sulphate: added as a manganese stabilizer and buffering agent.

Hydroxylamine sulphate: antioxidant.

- 5        The electrolysis process is carried out in diaphragm cells in which the anolyte and catholyte are separated by a semi-permeable material.

10       With regard to the flow of liquors in the separated electrolytes cell, the catholyte, when passing through the diaphragm, feeds the anode cell, such that the anolyte has a composition similar to the catholyte, although with a lower pH and noticeably manganese-depleted. The consumed electrolyte is therefore suitable for its recirculation to the process start. The cells must be kept at a certain temperature, and the compositions of the electrolytes must be kept homogenous.

- 15       Over time, the metal is deposited on the surface of the cathode in the form of flakes. The metal load is separated by mechanical means.

20       While the electrolytic manganese flakes are deposited on the cathode, manganese dioxide accumulates on the anode. Once the process has concluded, this product also requires washing with water and is subsequently separated by mechanical means.

#### Primary purification

- 25       It is carried out in the lixiviation reactor itself and is achieved by increasing the pH of the solution. To separate the waste contained in the resulting pulp, which has non-etched material as well as the precipitated impurities in the form of hydroxides (mainly Fe and Al), it is subjected to filter press filtration. The waste separated here requires washing with water to recover part of the manganese which it has entrained and to improve its chemical and physical features before its subsequent deposition.

The wash water from these filtration cakes can be used as addition water to the mixer.

The electrolytic manganese obtained by the process is in the form of flakes and has a Mn content of 99.9 %.

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## TESTS FOR OBTAINING MANGANESE SULPHATE LIQUOR

### EXAMPLE

10 1 kg of material with a humidity of 40% and an Mn content of 15% was used to start.

The material was mixed with 390 g of sulfuric acid and 390 ml of water in a ceramic recipient.

15 The mixture was then poured on a tray which was introduced in a furnace maintained at 300°C for 30 minutes.

20 The lixiviation was carried out with prepared synthetic anolyte. In this process, the residence time for the extraction of Mn was 1 hour, during which the pulp was maintained under strong stirring. After that time elapsed, 70 g of lime were added to the same lixiviation reactor and it was left under stirring for half an hour, the pH being increased from 3.7 to 6.5 in that time. To separate the waste contained in the resulting pulp, which has both non-etched material and the precipitated impurities in the form of hydroxides (mainly Fe and Al), it was subjected to vacuum filtration.

25 The waste separated here was subjected to washing with water to recover part of the manganese that it had entrained and to improve its chemical and physical characteristics prior to its subsequent deposition. According to the characterization carried out, this waste is of an inert type according to the applicable standards.

To remove from the liquor separated in the filtration the content it may have both of organic material as well as of traces of contaminants, it was passed through an active carbon filter.

5                      The second purification phase involved adding 11.1 cc of sulphide and 0.65 g of zinc sulphide primer. The obtained precipitate was separated by means of filtration. 1.25 l of liquor were thus obtained, with the following chemical analysis:

Mn (g/l)	45
Zn (ppm)	15
Ca (ppm)	250
Mg (g/l)	3,5
Fe (ppm)	<1

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## ELECTROLYTIC MANGANESE APPLICATIONS

15                      Electrolytic manganese is basically used in the aluminum industry, the resistance and ductility it provides (structural applications, resistant thin sheets, aeronautics, canning...), being capable of being supplied: in the parent alloy, as an injectable dust and in mixed dust briquettes (or mechanical pre-alloy). Other uses of electrolytic manganese are: in the steel industry as a desulphurizing and fine alloying agent for high performance stainless steels and HSLA steels; in the copper and nickel alloy industry; electrolytic manganese flakes as a chemical reaction catalyst; manganite  
20                      production for the manufacture of temperature-controlled variable resistors; the



manufacture of zinc-manganese ferrites for intermediate power electronic applications and as a pigment; manufacture of welding rods.

## PROCESS DIAGRAMS

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Description of the Diagrams of a Preferred Embodiment of the Invention.

Figure 1 shows the block diagram of the process which starts with the sulphation phase (1), followed by the hydrometallurgical phase consisting of four steps:  
10 lixiviation (3), primary purification (4), secondary purification (5) and conditioning (6), to finally reach the electrolysis (7) phase for obtaining electrolytic manganese.

Figure 2 schematically shows the diagram corresponding to the sulphation phase (1) involving the processing in a furnace (8) of the material supplied by a mixer (9),  
15 gasses being generated which are extracted in a forced manner from said furnace (8) through a chimney (10) and neutralized in a washing column (11) with the suitable reagent.

Figure 3 schematically shows the lixiviation (3) and primary purification (4) steps in  
20 which the product resulting from the sulphation phase (1) is treated in a tank (12) with an antacid coating, a pulp being obtained which is passed through a filter press (13) where washing of the filtration cakes can be further carried out countercurrently. This washing is carried out with water supplied from a tank (14), which can subsequently be used both in the initial mixture of the process and in the lixiviation  
25 phase. The inert waste is collected in the pan (15) at the bottom of the filter press.

Figure 4 corresponds to the secondary purification step (5), in which the primary liquor of the previous step is filtered with an active carbon filter (16) prior to being taken to the secondary purification tank (17), where the remaining impurities

precipitate in the form of sulphides which are separated through another filtration process (18), purified liquor (19) and a waste product, ZnS (20), finally being obtained.

5        Figure 5 schematically shows the conditioning step (6) in which the purified liquor is subjected to a pH increase in a tank (21), to be subsequently treated in a crystallizer (22) where the precipitation of Ca and Mg ammonium salts (24) of the liquor is favored.

10       Figure 6 schematically shows the electrolysis phase (7) of the liquor which is treated in a cell (26) provided with a catholyte heated by the action of a heat exchanger existing in the feed tank (25). The cell is a polyester diaphragm cell, and two anodes and one cathode are immersed inside, the latter being inside its own compartment. There is an immersion tank (27) to carry out the washing of the cathode after the  
15       deposition of the electrolytic Mn on the surface thereof. On the other hand, the anode sludges of the cell are accumulated and evacuated from the double bottom of the cell, passing to be treated in a sludge flocculation tank (28).

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